

## Additions and Corrections

---

**Estimation of Solvolysis Rates of Cyclic Secondary Substrates** [*J. Am. Chem. Soc.*, **96**, 4693 (1974)]. By J. MILTON HARRIS\* and SAMUEL P. MCMANUS, Department of Chemistry, The University of Alabama in Huntsville, Huntsville, Alabama 35807.

In Table I, the  $\sigma_t^*$  value for 7-norbornadienyl (7) should read -1.12.

**Effect of Crown Ether on the Conformational Equilibrium of Sodium Acetylacetonate** [*J. Am. Chem. Soc.*, **96**, 6184 (1974)]. By ERIC A. NOE and MORTON RABAN,\* Department of Chemistry, Wayne State University, Detroit, Michigan 48202.

Further nmr experiments have demonstrated that the two configurations of the acetylacetonate anion (1) present in pyridine solution containing crown ether are *E,Z*-1 and *Z,Z*-1 rather than *E,E*-1 and *Z,Z*-1 as previously reported. The *E,Z*-1 isomer gives rise to two equally intense singlets. However, the low-field singlet overlaps with the broad, intense resonance due to the crown ether, and we incorrectly concluded that the resonance at ca.  $\delta$  2.38 was the single resonance from the homotopic methyl groups of *E,E*-1 rather than one of a pair of singlets due to the diastereotopic methyl groups of *E,Z*-1. As a result of this change in configurational assignment of the downfield isomer, the isomer ratios reported on p 6186 must also be changed. The relevant sentences should read: "A temperature increase of only 6° lowers the percentage of *E,Z*-1 from 46 to 37% (Figures 1a and 1b). In another experiment with the same crown/enolate ratio, the following percentages were observed as a function of temperature: -46°, 40%; -51°, 48%; -60°, 54%; -65°, 58%."

**Matrix Photolysis of 1,2,3-Thiadiazole. On the Possible Involvement of Thiirene** [*J. Am. Chem. Soc.*, **96**, 6768 (1974)]. By A. KRANTZ\* and J. LAURENI, Department of Chemistry, State University of New York, Stony Brook, New York 11794.

The sentence beginning on the sixth line in the left-hand column on page 6769 should read, "The assignment of these bands to a single species is most clearly evident upon irradiation with light of  $\lambda > 220$  nm, which results in loss of the thioketene spectrum and the appearance of a band at 1520  $\text{cm}^{-1}$ , perhaps due to a secondary product from carbon monosulfide."

**Remote Anodic Substitution of Ketones** [*J. Am. Chem. Soc.*, **97**, 853 (1975)]. By JAMES Y. BECKER, LARRY R. BYRD, LARRY L. MILLER,\* and YING-HUNG SO, Department of Chemistry, Colorado State University, Fort Collins, Colorado 80521.

Two entries in Table II were interchanged. The first entry under "Product" should be 4-acetamido-4-methyl-2-hexanone and the last entry in this column should be 5-acetamido-2-hexanone.

**Enantioselective Oxirane Synthesis by Means of Dimethylsulfonium Methylide and Chiral Phase-Transfer Catalysts** [*J. Am. Chem. Soc.*, **97**, 1626 (1975)]. By TAMEJIRO HIYAMA,\* TAKASHI MISHIMA, HIROYUKI SAWADA, and HITOSI NOZAKI, Department of Industrial Chemistry, Kyoto University, Yoshida, Kyoto 606, Japan.

To the authors' regret the samples of the reported product, 2-phenyloxirane (1), were contaminated with 2-methyl-3-phenyloxirane ( $[\alpha]^{25D} 71.1^\circ$  (*c* 0.94),  $54^\circ$  (*c* 0.20) in acetone) derived from the phase-transfer catalysts. The rotational data are being reexamined, and further work to determine the degree of asymmetric induction is now in progress.

**Phase Transfer Catalysis. An Evaluation of Catalysts** [*J. Am. Chem. Soc.*, **97**, 2345 (1975)]. By ARTHUR W. HERRIOTT\* and DONALD PICKER, Department of Physical Sciences, Florida International University, Miami, Florida 33144, and the Department of Chemistry, State University of New York at Albany, Albany, New York 12222.

Footnote 31 should read: "We thank General Mills, Kankakee, Ill., for a sample of trioctylmethylammonium chloride, distributed as Aliquat 336."

**Irreversible Inhibition of  $\Delta^5$ -3-Ketosteroid Isomerase by 5,10-Secosteroids** [*J. Am. Chem. Soc.*, **97**, 2576 (1975)]. By F. H. BATZOLD and C. H. ROBINSON,\* The Johns Hopkins University School of Medicine, Baltimore, Maryland 21205.

On page 2577, second column, 25 lines down, the equation

$$\ln \frac{(\epsilon)}{[E_t]} = \frac{-k_3}{1 + (K_1/[I])}$$

should read:

$$\ln \frac{(\epsilon)}{[E_t]} = \frac{-k_3 t}{1 + (K_1/[I])}$$

Reference 14 should read, "R. Kitz and I. B. Wilson, *J. Biol. Chem.*, **237**, 3245 (1962)."

**A General Synthetic Method for Non-K-Region Arene Oxides** [*J. Am. Chem. Soc.*, **97**, 3185 (1975)]. By HARUHIKO YAGI and DONALD M. JERINA,\* National Institute of Arthritis, Metabolism and Digestive Diseases, National Institutes of Health, Bethesda, Maryland 20014.

On page 3189 in the preparation of 3,4-dihydrophenanthrene (8a), the oil on standing crystallized, mp 61-63°. On the same page, the melting point of 1,2-dibromo-1,2,3,4-tetrahydrophenanthrene (9a) (81-83°) is incorrect. The correct value is mp 116-117°.

On page 3191, the names of compounds 20c, 20d, and 21a should be corrected as follows: 8-bromo-7-trifluoroacetoxy-7,8,9,10-tetrahydrobenzo[a]pyrene (20c), 7-acetoxy-8-bromo-7,8,9,10-tetrahydrobenzo[a]pyrene (20d), and 10-acetoxy-7,9-dibromo-7,8,9,10-tetrahydrobenzo[a]pyrene (21a).